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(71) Applicant:

**Mitsui Mining & Smelting Co., Ltd.
Tokyo 141-8584 (JP)**

(72) Inventors:

- Numata, Koichi
Takehara-shi, Hiroshima 725-0024 (JP)
- Shinkawa, Hiromu
Toyota-gun, Hiroshima 729-2403 (JP)

(74) Representative:

**Forstmeyer, Dietmar, Dr. rer. nat., Dipl.-Chem. et
al
Boeters & Bauer,
Bereiteranger 15
81541 München (DE)**

(54) METHOD FOR PRODUCING LITHIUM MANGANATE

(57) A process for producing lithium manganate which is characterized by comprising pulverizing electrodeposited manganese dioxide having been crushed and neutralized in a pulverizer of the system in which particles are made to collide with each other, screening the grinds, mixing the resulting electrolytic manganese dioxide having an average particle size of 3 to 20 μm with a lithium raw material, and firing the mixture.

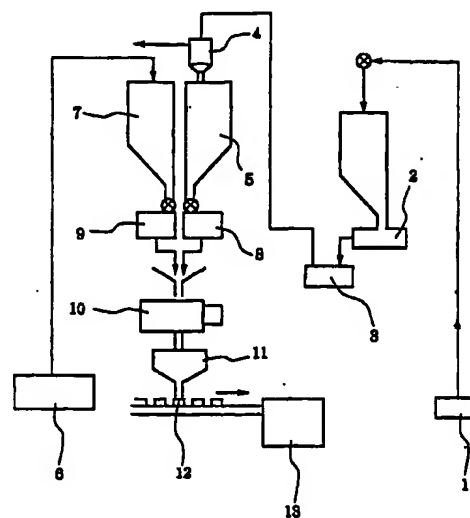


Fig. 1

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Description

Technical Field

[0001] The present invention relates to a process of producing lithium manganate. More particularly it relates to a process of lithium manganate production which hardly causes particles to be sintered during firing, reduces iron incorporation, and provides lithium manganate exhibiting satisfactory battery characteristics when used as a cathode material of a nonaqueous secondary battery.

Background Art

[0002] With the recent rapid development of portable and wireless electronic equipment such as personal computers and telephones, the demand for secondary batteries as a driving power source has been increasing. In particular nonaqueous secondary batteries are expected for their smallest size and high energy density. Cathode materials for nonaqueous secondary batteries meeting the demand include lithium cobaltate (LiCoO_2), lithium nickelate (LiNiO_2), lithium manganate (LiMn_2O_4), etc. Having an electrode potential of 4 V or higher with respect to lithium, these complex oxides are capable of providing batteries having a high energy density.

[0003] Of the above-described complex oxides, LiCoO_2 and LiNiO_2 have a theoretical capacity of about 280 mAh/g, while LiMn_2O_4 has a theoretical capacity as low as 148 mAh/g but is deemed suited for use in electric vehicles and the like because of an abundant and inexpensive supply of manganese oxide as a raw material and freedom from such thermal instability in charging as observed with LiNiO_2 .

[0004] Electrolytic manganese dioxide is a suitable manganese source of lithium manganate (LiMn_2O_4) for its inexpensiveness and abundance. Electrodeposited manganese dioxide is usually crushed and neutralized for use as a manganese raw material.

[0005] The crushed and neutralized electrolytic manganese dioxide has an average particle size of about 20 to 40 μm . It should be pulverized to an average size of about 5 μm which is suitable as a manganese raw material. However, the electrolytic manganese dioxide obtained by conventional pulverization has an average particle size of about 10 μm . Where lithium manganate prepared from such electrolytic manganese dioxide is used as a cathode material for nonaqueous secondary batteries, there arise problems that the coating properties are poor, and high-rate characteristics are impaired. Moreover, there is another problem that impurities such as iron may be incorporated into the electrolytic manganese dioxide, which adversely affects the properties of lithium manganate.

[0006] The pulverization is accompanied by generation of fine particles of the electrolytic manganese

dioxide, which, if used as a manganese raw material, tends to accelerate sintering among particles while fired with a lithium raw material.

[0007] Accordingly, it is an object of the present invention to provide a process for producing lithium manganate, which hardly causes particles to be sintered during firing, reduces iron incorporation, and provides lithium manganate exhibiting satisfactory battery characteristics when used as a cathode material of a nonaqueous secondary battery.

Disclosure of the Invention

[0008] As a result of studies, the present inventors have found that the above object is accomplished by pulverizing electrolytic manganese dioxide in a pulverizer of the system in which particles are made to collide with each other into particles of about 5 μm as a raw material of lithium manganate and recovering fine particles by-produced in the pulverization.

[0009] The present invention has been completed based on the above finding and provides a process of producing lithium manganate which is characterized by comprising pulverizing electrodeposited manganese dioxide having been crushed and neutralized in a pulverizer of the system in which particles are made to collide with each other, screening the grinds, mixing the resulting electrolytic manganese dioxide having an average particle size of 3 to 20 μm with a lithium raw material, and firing the mixture.

Brief Description of the Drawing

[0010]

Fig. 1 schematically illustrates an embodiment of the production process according to the present invention.

Best Mode for Carrying out the Invention

[0011] The production process of the present invention will be described in detail with reference to the drawing. Fig. 1 is a schematic illustration showing an embodiment of the process according to the invention, wherein numeral 1 is an electrolytic manganese dioxide tank; 2, a constant powder feeder; 3, a pulverizer (a pulverizer of the system to make particles collide with each other); 4, a screen classifier; 5, a raw material tank (A); 6, a lithium raw material; 7, a raw material tank (B); 8, a metering tank (A); 9, a metering tank (B); 10, a mixer; 11, a hopper; 12, a firing container; and 13, a firing furnace.

[0012] The electrolytic manganese dioxide which is a manganese raw material of lithium manganate is produced electrolytically in the electrolytic cell. Electrolysis is carried out by, for example, using a manganese sulfate solution having a prescribed concentration as an

electrolytic solution, a carbon plate as a cathode, and a titanium plate as an anode. Electrolysis is conducted at a constant current density while heating to electrodeposit manganese dioxide on the anode. The electrodeposited manganese dioxide is peeled off the anode, crushed, and neutralized with sodium hydroxide, etc. The crushed and neutralized electrolytic manganese dioxide has an average particle size of about 20 to 40 μm .

[0013] As shown in Fig. 1, the electrolytic manganese dioxide having an average particle size of 20 to 40 μm that is stored in the electrolytic manganese dioxide tank 1 is fed to the pulverizer 3 through the constant powder feeder 2 and pulverized in the pulverizer 3. The pulverizer 3 used in the invention is of the system in which particles are made to collide with each other. PJM SYSTEM JET MILL, manufactured by Nippon Pneumatic Kogyo K.K., can be mentioned as an example. By use of such a pulverizer, the average particle size of the electrolytic manganese dioxide is reduced to 20 μm or smaller.

[0014] The electrolytic manganese dioxide obtained by pulverizing by means of conventionally employed mills has an average particle size of about 15 μm . Where such electrolytic manganese dioxide is used as manganese source of lithium manganate, the resulting lithium manganate is, when used as a cathode material of nonaqueous secondary batteries, liable to have poor coating properties and to impair the high-rate characteristics. In addition, there arises another problem that iron may be incorporated into the electrolytic manganese dioxide, which adversely affects the properties of lithium manganate. Where the above-described pulverizer of the system in which particles are made to collide with each other is used, the problem of iron incorporation into electrolytic manganese dioxide does not occur.

[0015] The electrolytic manganese dioxide having been pulverized in the pulverizer 3 of the system in which particles are made to collide with each other is sent to the screen classifier 4, where it is screened. The screened electrolytic manganese dioxide having an average particle size of 3 to 20 μm , e.g., 5 μm , is reserved in the raw material tank (A).

[0016] On the other hand, the undersize electrolytic manganese dioxide having an average particle size smaller than 3 μm is discharged out of the system and recovered by means of a cyclone, etc. for another use. The electrolytic manganese dioxide whose average particle size is smaller than 3 μm is produced through pulverization in a proportion of about 5 to 20% by weight based on the total weight of the electrolytic manganese dioxide. In the present invention, since such fine particles of electrolytic manganese dioxide are not used as a manganese raw material for lithium manganate, sintering among particles which would be caused by such fine particles on firing can be prevented.

[0017] The lithium raw material 6 to be used

includes lithium carbonate (Li_2CO_3), lithium nitrate (LiNO_3), and lithium hydroxide (LiOH). Lithium carbonate is particularly preferred. The lithium raw material is reserved in the raw material tank (B) 7.

[0018] Prescribed amounts of the electrolytic manganese dioxide reserved in the raw material tank (A) 5 and the lithium raw material, such as lithium carbonate, reserved in the raw material tank (B) 6 are weighed out in the metering tank (A) 8 and the metering tank (B) 9, respectively, and mixed together in the mixer 10. The molar ratio of Li in the lithium raw material to Mn in the electrolytic manganese dioxide, Li/Mn , is preferably 0.50 to 0.60.

[0019] The mixed raw material from the mixer 10 is packed into the firing container 12 through the hopper 11. The mixed raw material can be used either as such or after being granulated. Granulation can be carried out in either a wet system or a dry system. Methods of granulation include piston granulation, tumbling granulation, fluidized bed granulation, mixing granulation, spray drying, pressure forming granulation, and flaking granulation using a roll, etc.

[0020] The mixed raw material packed in the firing container 12 is put in the firing furnace 13 and fired at 600 to 1000°C, preferably 750 to 1000°C, to obtain lithium manganate, for example, spinel-type lithium manganate. The firing furnaces which can be used include a pusher-type kiln, a roller hearth kiln, a rotary kiln, and a stationary furnace. The firing time is 1 hour or longer, preferably 5 to 20 hours. During the firing the above-mentioned sintering among particles, which is attributed to fine manganese dioxide particles, hardly takes place.

[0021] In this manner, lithium manganate can be obtained. The resulting lithium manganate suffers little incorporation of iron as stated above. The lithium manganate thus obtained has satisfactory coating properties and is fit for use as a cathode material of nonaqueous secondary batteries.

[0022] In the nonaqueous secondary battery of the present invention, the above-described cathode material is mixed with a conductive material, such as carbon black, and a binder, such as Teflon binder, to prepare a cathode active material mixture. For the anode, lithium or a material capable of intercalating and disintercalating lithium, such as carbon, is used. While not limiting, nonaqueous electrolytes which can be used include a lithium salt, e.g., lithium hexafluorophosphate (LiPF_6), dissolved in a mixed solvent, such as ethylene carbonate/dimethyl carbonate.

[0023] The nonaqueous secondary battery according to the present invention exhibits improved battery characteristics, such as high-rate characteristics.

Industrial Applicability

[0024] According to the production process of the present invention, sintering of particles hardly occurs during firing, incorporation of iron is inhibited, and lith-

ium manganate exhibiting satisfactory battery characteristics when used as a cathode material of a nonaqueous secondary battery can be obtained.

Claims

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1. A process for producing lithium manganate which is characterized by comprising pulverizing electrode-
posited manganese dioxide having been crushed
and neutralized in a pulverizer of the system in 10
which particles are made to collide with each other,
screening the grinds, mixing the resulting electro-
lytic manganese dioxide having an average particle
size of 3 to 20 μm with a lithium raw material, and
firing the mixture. 15
2. A process for producing lithium manganate as set
forth in claim 1, wherein the screen out electrolytic
manganese dioxide having an average particle size
smaller than 3 μm is recovered. 20
3. A process for producing lithium manganate as set
forth in claim 1, wherein the firing is carried out at
750°C or higher. 25
4. A cathode material for a nonaqueous secondary
battery which comprises lithium manganate
obtained by the process set forth in claim 1.
5. A nonaqueous secondary battery composed of a 30
cathode comprising the cathode material set forth
in claim 4, an anode capable of intercalating and
disintercalating lithium, and a nonaqueous electro-
lytic solution. 35

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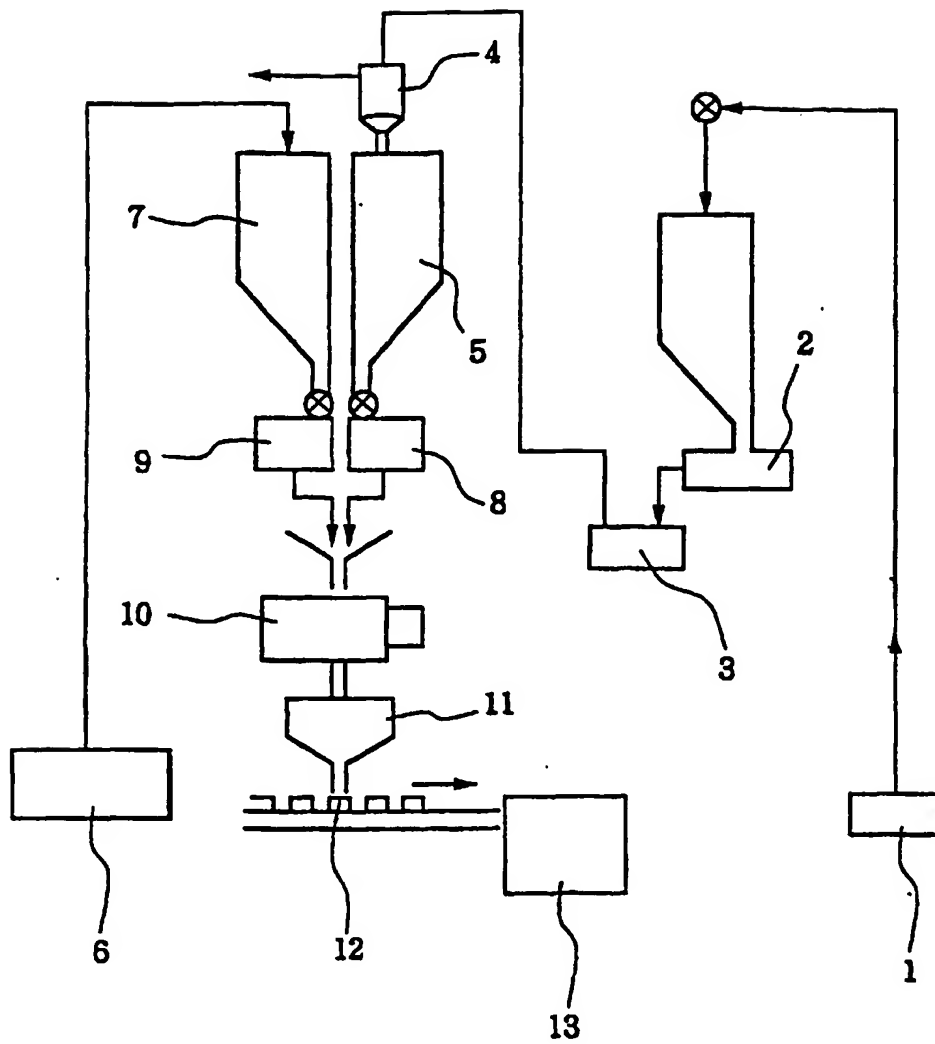


Fig. 1

INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP99/04811

A. CLASSIFICATION OF SUBJECT MATTER Int.Cl. ⁶ C01G45/00		
According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED Minimum documentation searched (classification system followed by classification symbols) Int.Cl. ⁶ C01G45/00, H01M4/58, H01M4/02		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched Jitsuyo Shinan Koho 1926-1996 Toroku Jitsuyo Shinan Koho 1994-1999 Kokai Jitsuyo Shinan Koho 1971-1999 Jitsuyo Shinan Toroku Koho 1996-1999		
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
EX	JP, 11-126607, A (Japan Metals & Chemicals Co., Ltd.), 11 May, 1999 (11.05.99), Claims; Par. Nos. [0021] to [0024]; example (Family: none)	1-5
X	JP, 09-213333, A (Japan Metals & Chemicals Co., Ltd.), 15 August, 1997 (15.08.97), Claims; example (Family: none)	1-5
X	JP, 06-295724, A (Mitsui Mining & Smelting Co., Ltd.), 21 October, 1994 (21.10.94), Claims; Par. Nos. [0009], [0014]; example (Family: none)	1-5
A	JP, 09-063583, A (Toshiba Battery Co., Ltd.), 18 March, 1997 (18.03.97), Claims; Par. Nos. [0010] - [0011]; example (Family: none)	1-5
A	JP, 05-174823, A (Hitachi Maxell, Ltd.), 13 July, 1993 (13.07.93), Claims; example (Family: none)	1-5
A	JP, 10-265224, A (Mitsubishi Chemical Corporation),	1-5
<input checked="" type="checkbox"/> Further documents are listed in the continuation of Box C. <input type="checkbox"/> See patent family annex.		
* Special categories of cited documents: "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family		
Date of the actual completion of the international search 08 December, 1999 (08.12.99)		Date of mailing of the international search report 21 December, 1999 (21.12.99)
Name and mailing address of the ISA/ Japanese Patent Office		Authorized officer
Facsimile No.		Telephone No.

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INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP99/04811

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	06 October, 1998 (06.10.98), Claims; example (Family: none) JP, 08-295516, A (Mori Energy 1990 Ltd.), 12 November, 1996 (12.11.96), Claims; example & DE, 19615686, A1 & GB, 229988, A & FR, 2733223, A1 & CA, 2147578, A & US, 5882218, A	1-5

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